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Chapter 11 Chemical Reactions- Chemistry by Ms.Basima Pearson  
Chemistry Chapter 11: Section 2: Types of Chemical Reactions  

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The chemical reactions that occur inside the human body are extremely complex. One of the main reactions that occurs is the metabolism of glucose, and another reaction is the detoxification of the liver. The metabolism of glucose is one of the main factors that affect the levels of insulin that are generated by the liver.

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Section 11.1 Assessment. Describe the steps in writing a balanced chemical equation. Write the skeleton equation for the following reactions: Heating solid copper(II)sulfide in the presence of oxygen gas produces pure copper and sulfur dioxide gas. Iron metal and chlorine gas react to form solid iron(III)chloride.  $\text{CuS (s)} + \text{O}_2\text{(g)} \rightarrow \text{Cu (s)} + \text{SO}_2\text{(g)}$  D

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330 Chapter 11 11.2 Types of Chemical Reactions Often charcoal briquettes provide the heat for barbeque grills through the burning of carbon. Have you ever felt the heat and smelled the smoke coming from a burning charcoal grill? The heat and smoke are the products of a combustion reaction. Combustion is one of the five general types of chemical

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oxidation – reduction reactions A chemical reaction that exhibits a change in the oxidation states of one or more elements in the reactants that has the general form oxidant + reductant → reduced oxidant + oxidized reductant. The general forms of these five kinds of reactions are summarized in Table 11.6.1, along with examples of each.

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As shown in Figure 11.3.1, applying a small amount of heat to a pile of orange ammonium dichromate powder results in a vigorous reaction known as the ammonium dichromate volcano. Heat, light, and gas are produced as a large pile of fluffy green chromium(III) oxide forms. We can describe this reaction with a chemical equation An expression that

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gives the identities and quantities of the ...

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11.2 Types of Chemical Reactions > 30 A halogen can also replace another halogen from a compound. • The activity of halogens decreases as you go down Group 7A of the periodic table—fluorine, chlorine, bromine, and iodine. • Bromine is more active than iodine, so this reaction occurs:  $\text{Br}_2(\text{aq}) + 2\text{NaI}(\text{aq}) \rightarrow 2\text{NaBr}(\text{aq}) + \text{I}_2(\text{aq})$

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