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(changes in pressure) SL

7.1 Le Chatelier's

principle (changes in

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Temperature Changes

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~~Chatelier's principle~~

~~(concentration and~~

~~pressure) SL Which way~~

~~will the Equilibrium~~

~~Shift? (Le Chatelier's~~

~~Principle)~~

Lab Experiment #13:

The Equilibrium

Constant.

Equilibrium and Cobalt

Complex Ions |

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Chatelier's principle has
very important and wide
application in chemical
reactions that are both
exothermic and
endothermic in nature. It
aims at maintaining an

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equilibrium position for all reacting species even one any factor is altered so that the

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Answers altered so that
the reaction can proceed
to completion.

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Chatelier's Principle In

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Chemical ... 3d.

According to Le Chatelier's Principle, predict which way should the reaction shift (left or right). Now to test your prediction, to test tube #2 add a few crystals

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The le Chatelier's
principle can be applied
to understand the effect
of change in pressure on
the systems at
equilibrium as follows.

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1) When the partial pressure of any of the gaseous reactants or of the products is increased, the position of equilibrium is shifted so as to decrease its partial pressure .

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According to Le
Chatelier's Principle,
predict which way
should the reaction shift
(left or right). Now to
test your prediction, to
test tube #2 add a
few crystals of solid
potassium thiocyanate

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(KSCN). 3e. What do you observe happening? Record any color changes or other changes you observe. 3f.

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principle has very
important and wide
application in chemical
reactions that are both
exothermic and
endothermic in nature. It

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aims at maintaining an equilibrium position for all reacting species even one any factor is altered so that the reaction can proceed to completion.

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The effect of concentration on a system at equilibrium depends on whether the change in concentration is affecting a reactant or product species. In general when the concentration of a species is increased, the system will

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PRINCIPLE PART II

In this part of the experiment you will investigate the



Procedure: 1.

Examine solutions of $\text{K}_2\text{Cr}_2\text{O}_7$ and K_2CrO_4 .

Fill in Table III. 2. Put three drops of K_2CrO_4 into a well. Add three drops HNO_3 . Record any color change.

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Record your
observations in Table
IV. 3.

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