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Equilibrium, Cu<sup>2+</sup> and NH<sub>3</sub> complex ion Fluid, Electrolyte, and Acid Base Balance

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ISBN: 0867201487 9780867201482: OCLC Number: 21346795: Notes: "A second edition based on the earlier book, Mineral equilibria, written by Dr. Garrels and published in 1960, in the Harper's geoscience series"--Page vii.

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ous systems involving salt-type minerals are examined in this work. Stability relations have been discussed on the basis of bulk equilibria. MINERAL-SOLUTION EQUILIBRIA The equilibria in selected salt-type mineral systems with special reference to calcite and apatite are examined below. Theoretical results are correlated with experimental data.

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The observed changes in the electrokinetic properties of calcite and apatite can be examined on the basis of the mineral-solution chemical equilibria involving dissolved mineral species. From studies of solubility isotherms for apatite and calcite at 25°C (Fig. 3.4, Fig. 3.5), the singular point for these minerals is identified to be 9.3. Above this pH calcite is more stable than apatite.

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3. SOLUTION-MINERAL EQUILIBRIA PART 1: CARBONATES. Carbonic acid and the carbonate minerals provide another good illustration of the use of equilibrium reasoning in geochemistry. Interactions among these compounds determine the conditions under which limestones and dolomites are formed or dissolved, and likewise the conditions of formation of carbonate minerals as cements in soils and sandstones and as vein fillings.

[SOLUTION-MINERAL EQUILIBRIA PART 1: CARBONATES](#)

sulted in the final version, Solutions, Minerals, and Equilibria, about twice as long, in 1965. The original edition was published by Harper; Harper was taken over by Harper & Row, which permitted the book to go out of print. Fortunately, W.H. Freeman of Freeman, Cooper & Company re-printed the book, 1 and since 1975, it has been selling successfully.

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